
CHEMISTRY

9701/35

Paper 3 Advanced Practical Skills 1

May/June 2017

MARK SCHEME

Maximum Mark: 40

Published

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Question	Answer	Marks
1(a)	I Constructs a table for results showing volume of FA 1 , volume of water, reaction time, reaction rate for all experiments carried out	1
	II Appropriate headings and units for recorded data given. Volumes in cm ³ or / cm ³ or (cm ³). Time in seconds or / s or (s) All volumes except zero given to .00.	1
	III All times recorded to the nearest second.	1
	IV 3 additional volumes chosen intervals not less than 2.00 cm ³ and all volumes of FA 1 ≥ 6.00 cm ³ and one volume of FA 1 ≤ 8.00 cm ³	1
	V In all 3 additional experiments water is added to make a total of 20.(00) cm ³	1
	VI + VII Compare time for 20.00 cm ³ of FA 1 with that of supervisor. 2 marks for ± 3 s 1 mark for ± 5 s	2
	VIII Compare ratio of time for 10.00 cm ³ of FA 1 / time for 20.00 cm ³ of FA 1 . 1 mark for ratio between 1.8 – 2.2	1
	IX All rates correctly calculated using 500 / time (minimum 2 sf and 1 dp)	1
	X Units for rate given as s ⁻¹	1

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Question	Answer	Marks
1(b)	I Rate on <i>y-axis</i> and volume on <i>x-axis</i> . Axes clearly labelled and suitable linear scales.	1
	II Scale chosen to use more than half of each axis for origin and plotted points	1
	III All points plotted correctly to within half a square and in the correct square.	1
	IV Draws a line of best fit. This may be a straight line or a smooth curve with anomalous points indicated.	1
1(c)	Rate is (directly) proportional to concentration of peroxodisulfate or comment suitable to shape of graph	1
1(d)(i)	Reads rate from graph correct to one small square and shows use of this number in calculation	1
	Shows use of $500 / \text{rate}$	1
1(d)(ii)	Correctly calculates $(0.5 / \text{time for expt 1}) \times 100$ to 2 or more sf	1
1(d)(iii)	The student is correct as the reaction time would be longer and so the (percentage) error reduced.	1
1(d)(iv)	There is so much thiosulfate that all the iodide reacts so there is no iodine to turn the starch blue-black.	1

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Question	Answer	Marks
1(e)(i)	Record time to nearest second with units of s	1
	Candidate's time compared with that from Expt 1. 1 mark for ± 3 s	1
1(e)(ii)	Estimates a time as $4 \times$ ans (i)	1
	Time / rate related to concentration of $\text{S}_2\text{O}_3^{2-}$ / FA 3 Increased concentration of FA 3 increases time of reaction / time longer / decreases rate of reaction / rate lower / smaller / reaction slower.	1
	Total:	24

Question	Answer	Marks																								
FA 4 is $(\text{NH}_4)_2\text{Fe}(\text{SO}_4)_2$ FA 5 is $\text{KAl}(\text{SO}_4)_2$ FA 6 is Na_2SO_3 FA 7 is H_2SO_4 FA 8 is NaNO_2																										
2(a)(i)	<table border="1" data-bbox="427 331 1563 938"> <thead> <tr> <th data-bbox="427 331 577 459" rowspan="2"><i>test</i></th> <th colspan="2" data-bbox="582 331 1411 395"><i>observation</i></th> <th data-bbox="1415 331 1563 459" rowspan="2"><i>mark</i></th> </tr> <tr> <th data-bbox="582 399 992 459">FA 4</th> <th data-bbox="996 399 1411 459">FA 5</th> </tr> </thead> <tbody> <tr> <td data-bbox="427 462 577 598" rowspan="2">+ NaOH</td> <td data-bbox="582 462 992 526">green ppt</td> <td data-bbox="996 462 1411 526">white ppt</td> <td data-bbox="1415 462 1563 526">1</td> </tr> <tr> <td data-bbox="582 529 992 598">insoluble in excess</td> <td data-bbox="996 529 1411 598">soluble in excess</td> <td data-bbox="1415 529 1563 598">1</td> </tr> <tr> <td data-bbox="427 601 577 702">then warm</td> <td data-bbox="582 601 992 702">gas / ammonia turns (damp red) litmus blue</td> <td data-bbox="996 601 1411 702">no reaction / litmus stays red</td> <td data-bbox="1415 601 1563 702">1</td> </tr> <tr> <td data-bbox="427 705 577 869" rowspan="2">+ NH₃</td> <td data-bbox="582 705 992 869">green ppt and turning brown (in air) in either alkali test</td> <td data-bbox="996 705 1411 869">white ppt</td> <td data-bbox="1415 705 1563 869">1</td> </tr> <tr> <td data-bbox="582 873 992 938">insoluble in excess</td> <td data-bbox="996 873 1411 938">insoluble in excess</td> <td data-bbox="1415 873 1563 938">1</td> </tr> </tbody> </table>	<i>test</i>	<i>observation</i>		<i>mark</i>	FA 4	FA 5	+ NaOH	green ppt	white ppt	1	insoluble in excess	soluble in excess	1	then warm	gas / ammonia turns (damp red) litmus blue	no reaction / litmus stays red	1	+ NH ₃	green ppt and turning brown (in air) in either alkali test	white ppt	1	insoluble in excess	insoluble in excess	1	5
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Question	Answer	Marks
2(a)(ii)	FA 4 contains NH_4^+ and Fe^{2+} FA 5 contains Al^{3+} 2 marks for all three correct 1 mark for any two correct	2
2(b)	Selects $\text{BaCl}_2(\text{aq})$ or $\text{Ba}(\text{NO}_3)_2(\text{aq})$ followed by appropriate acid (acid must be named) OR Selects acidified potassium manganate(VII) OR Selects named acid and tests gas with acidified potassium manganate(VII)	1
	White ppt that is soluble in acid OR Decolourises (potassium manganate(VII))	1
	SO_3^{2-}	1
2(c)(i)	+ Mg Effervescence / fizzing / bubbles	1
	Gas / H_2 / fizz pops with a lighted splint	1
	+ FA 8 Brown (yellow / orange) fumes or gas turns blue litmus red/bleached or blue solution	1
2(c)(ii)	H_2SO_4	1
	NaNO_2	1
2(c)(iii)	$\text{Mg}(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{H}_2(\text{g})$	1
	Total:	16