
CHEMISTRY

9701/31

Paper 3 Advanced Practical Skills 1

October/November 2016

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2016 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.

| | | | |
|--------|--|----------|-------|
| Page 3 | Mark Scheme | Syllabus | Paper |
| | Cambridge International AS/A Level – October/November 2016 | 9701 | 31 |

| Question | Answer | Marks |
|----------|---|-------|
| 2(a) | I Initial and final burette readings and volume added recorded for rough titre and accurate titre details tabulated. [minimum 2 × 2 'boxes' with relevant information] | 1 |
| | II Initial and final burette readings recorded and volume of FA 3 added recorded for each accurate titration. Headings and units correct for accurate titrations Headings: initial / final (burette) reading / volume or reading / volume at start / finish and volume / FA 3 added / used or titre [not difference / total] allow vol but not V and Units: (cm ³) or / cm ³ or in cm ³ [or cm ³ by every entry] | 1 |
| | III All accurate burette readings are recorded to the nearest 0.05 cm ³ Do not award this mark if: 50(.00) is used as an initial burette reading; more than one final burette reading is 50(.00); any burette reading is greater than 50(.0) | 1 |
| | IV Final uncorrected titre is within 0.10 cm ³ of any previous uncorrected accurate titre. <i>Do not include a reading if it is labelled rough.</i> <i>Do not award the mark if any accurate burette readings (apart from the initial zero) are given as integers.</i> | 1 |

| | | | |
|---------------|---|-----------------|--------------|
| Page 4 | Mark Scheme | Syllabus | Paper |
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| Question | Answer | Marks |
|-----------------|---|--------------------------|
| | <p>V, VI and VII Examiner rounds any accurate burette readings to the nearest 0.05 cm^3, checks subtractions and then selects the 'best' accurate titres using the hierarchy: identical titres; titres within 0.05 cm^3; titres within 0.1 cm^3; etc., to calculate mean correct to 0.01 cm^3.</p> <p>Examiner uses the best titre to calculate the ratio of acid remaining after reaction with calcium carbonate in Question 1 to this best titre from Question 2 for Supervisor and each candidate.</p> <p>The ratio of this value for the candidate is compared to the ratio of this value for the Supervisor and marks awarded as follows.</p> <p>Award V, VI and VII for 0.95 – 1.05 Award V and VI for 0.90 – 1.10 Award V for 0.80 – 1.20</p> | <p>3</p> <p>7</p> |

| | | | |
|---------------|---|-----------------|--------------|
| Page 6 | Mark Scheme | Syllabus | Paper |
| | Cambridge International AS/A Level – October/November 2016 | 9701 | 31 |

| Question | Answer | Marks |
|-----------------|--|--|
| 2(c)(vii) | Correctly uses $[(vi) \times 100]/(\text{mass in (a)})$ to a minimum of 2 sf | 1 5 |
| 2(d) | Question 1: % purity lower as loss of gas means fewer moles /less mass CaCO₃ Question 2: no change /% same as same amount of acid reacts / (amount) acid left is same | 1 1 1 1 4 max 3 |
| | Total | 16 |

| Question | Answer | Marks | | | | | | | | |
|---|--|--------------|-------------|---|--|-----------------------------|---|--|--------------------------------------|-------------------------|
| FA 5 is NaNO ₃ (s); FA 6 is CuCO ₃ (s); FA 7 is NaBr(aq) | | | | | | | | | | |
| 3(a)(i) | <table border="1" style="width: 100%;"> <thead> <tr> <th style="width: 50%;">FA 5</th> <th style="width: 50%;">FA 6</th> </tr> </thead> <tbody> <tr> <td>(goes to) colourless or yellow liquid/ solution</td> <td>(green) powder / solid (turns) black / black residue</td> </tr> <tr> <td>gas relights glowing splint</td> <td>or gas turns limewater milky / cloudy white / chalky / forms white ppt</td> </tr> <tr> <td>gas (turns) brown / brown gas or solution turns blue</td> <td>(pale) blue solution / liquid formed</td> </tr> </tbody> </table> | FA 5 | FA 6 | (goes to) colourless or yellow liquid/ solution | (green) powder / solid (turns) black / black residue | gas relights glowing splint | or gas turns limewater milky / cloudy white / chalky / forms white ppt | gas (turns) brown / brown gas or solution turns blue | (pale) blue solution / liquid formed | 1+1 1 1+1 |
| FA 5 | FA 6 | | | | | | | | | |
| (goes to) colourless or yellow liquid/ solution | (green) powder / solid (turns) black / black residue | | | | | | | | | |
| gas relights glowing splint | or gas turns limewater milky / cloudy white / chalky / forms white ppt | | | | | | | | | |
| gas (turns) brown / brown gas or solution turns blue | (pale) blue solution / liquid formed | | | | | | | | | |

| | | | |
|---------------|---|-----------------|--------------|
| Page 7 | Mark Scheme | Syllabus | Paper |
| | Cambridge International AS/A Level – October/November 2016 | 9701 | 31 |

| Question | Answer | Marks | | | | | | | | |
|--|--|--|-------------|---|--|--|---|---|--|---------------------|
| 3(a)(ii)–(iv) | <table border="1"> <thead> <tr> <th>FA 5</th> <th>FA 6</th> </tr> </thead> <tbody> <tr> <td>(iii) solid dissolves / colourless solution allow no reaction / no change / no effervescence</td> <td>effervescence / fizzing / bubbling and blue solution / liquid formed</td> </tr> <tr> <td>(iv) no reaction / no change / no ppt / remains colourless</td> <td>blue ppt and insoluble in excess</td> </tr> <tr> <td>(v) no reaction / no change / no ppt / remains colourless</td> <td>(pale) blue ppt and soluble in excess to give deep / dark blue (solution)</td> </tr> </tbody> </table> | FA 5 | FA 6 | (iii) solid dissolves / colourless solution allow no reaction / no change / no effervescence | effervescence / fizzing / bubbling and blue solution / liquid formed | (iv) no reaction / no change / no ppt / remains colourless | blue ppt and insoluble in excess | (v) no reaction / no change / no ppt / remains colourless | (pale) blue ppt and soluble in excess to give deep / dark blue (solution) | 1 1 1 |
| | FA 5 | FA 6 | | | | | | | | |
| | (iii) solid dissolves / colourless solution allow no reaction / no change / no effervescence | effervescence / fizzing / bubbling and blue solution / liquid formed | | | | | | | | |
| (iv) no reaction / no change / no ppt / remains colourless | blue ppt and insoluble in excess | | | | | | | | | |
| (v) no reaction / no change / no ppt / remains colourless | (pale) blue ppt and soluble in excess to give deep / dark blue (solution) | | | | | | | | | |
| 3(a)(v) | FA 5: cation unknown; anion nitrate / NO_3^- FA 6: cation Cu^{2+} / copper(II); anion carbonate / CO_3^{2-} 4 correct = 3 marks 3 correct = 2 marks 2 correct = 1 mark | 1 1 1 | | | | | | | | |
| 3(a)(vi) | $\text{CuCO}_3(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ | 1 | | | | | | | | |
| 3(b)(i) | Selects AgNO_3 and NH_3 Selects NaOH and Al and HCl / HNO_3 / H_2SO_4 | 1 1 | | | | | | | | |
| 3(b)(ii) | Clearly defined test observation conclusion sections FA 7 + AgNO_3 cream ppt partially soluble in NH_3 FA 7 is bromide / Br^- from cream ppt | 1 1 1 | | | | | | | | |
| | Total | 17 | | | | | | | | |