

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

**MARK SCHEME for the October/November 2011 question paper  
for the guidance of teachers**

**0620 CHEMISTRY**

**0620/62**

Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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- 1 (a) (i) water/H<sub>2</sub>O inserted into box (1) [1]  
(ii) two arrows underneath magnesium and wool (1) [1]
- (b) magnesium oxide (1) [1]
- (c) lighted splint (1) pops (1) [2]  
glowing splint pops = 1
- (d) highly/very exothermic reaction/high temperature reached/suck back of water/owtte (1) [1]
- 2 (a) Table of results  
volumes correct (3) -1 for each incorrect  
0, 17, 25, 40, 48, 54, 57 [3]
- (b) points plotted correctly (3) -1 for each incorrect  
smooth curve missing anomalous point (1) [4]
- (c) (i) at 2 min (1) [1]  
(ii) from graph  $\pm$  half small square 30 cm<sup>3</sup> (1) indication on grid (1) [2]
- (d) (i) decreases/slows down (1) **not** stops [1]  
(ii) hydrochloric acid used up/hydrochloric acid becomes less concentrated (1) [1]  
**not** reactants used
- (e) (i) sketch curve to left of original (1) **ignore** if level is above original [1]  
(ii) sketch curve to right and below original (1) [1]
- 3 (a) to speed up the reaction/owtte (1) **not** reacts easily [1]
- (b) excess cobalt carbonate/base used (1) [1]
- (c) metal could react/glass does not react/owtte (1) [1]
- (d) solid/cobalt chloride visible/no more fizzing/no more gas (CO<sub>2</sub>) produced (1) [1]  
**ignore** colour change
- (e) crystals forming (on glass rod/on edge) (1) [1]

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- (f) anhydrous cobalt chloride formed/water/steam removed/powder formed (1) turn blue (1) [2]
- 4 (a) Table of results for Experiments 1 and 2  
initial boxes completed correctly 0.0, 2.0 (1)
- (b) final boxes completed correctly 23.0, 48.0 (1)  
differences correct 23.0, 46.0 (1) **allow** ecf  
readings to 1 dp (1) [4]
- (c) to remove impurities/solution F/owtte (1) [1]
- (d) as an indicator/to show presence of iodine/owtte (1) [1]
- (e) (i) Experiment 2 (1) [1]  
(ii) Experiment 2 2x volume Experiment 1 [1]  
(iii) solution **F** more concentrated/stronger (1) **allow** converse  
2x as concentrated (2) [2]
- (f) half value from table result for Experiment 1, 11.5 (1)  
half volume of potassium iodate/iodine/ $\frac{23}{2}$  (1) [2]
- (g) (i) two sources of error (2)  
e.g. experiment only done once/using a measuring cylinder to measure iodate/  
acid going past end point/owtte [2]  
**ignore** reference to temperature or human error  
(ii) two meaningful improvements related to above (2)  
e.g. use a pipette/burette/add smaller volumes e.g. 0.5 cm<sup>3</sup>/repeat experiment [2]
- 5 (a) (i) blue (1) [1]  
(b) white (1) precipitate (1) [2]  
(c) (i) blue (1) precipitate (1) [2]  
(ii) blue precipitate (1) dissolves/solution (1) deep/royal blue (1) [3]  
(e) organic (1) hydrocarbon / flammable / fuel (1) [2]

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6 (a) litmus paper/pH paper (1)  
blue/8–10 (1)  
test for  $\text{NH}_4^+$  using NaOH = 0  
correct chemical test and result e.g.  $\text{Cu}^{2+}$  could score 2 marks [2]

(b) 25 cm<sup>3</sup> of Kleen Up in flask/beaker (1) **not** test-tube  
nitric acid in burette (1)  
add indicator (1) no indicator = max 2  
add/titrate acid (1)  
until neutral/owtte (1)  
note volume acid (1)  
calculate concentration (1) max [5]

**[Total: 60]**