CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International General Certificate of Secondary Education

MARK SCHEME for the October/November 2014 series

0620 CHEMISTRY

0620/62

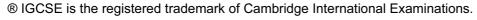
Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October / November 2014 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.





P	age 2	Mark Scheme	Syllabus	Paper
		Cambridge IGCSE – October / November 2014	0620	62
1	(a) (i) U-tube (1) gas jar (1) not: measuring cylinder		[2]
	(1	arrow inserted under shaded solid mixture (1)		[1]
		ess dense / lighter than air (1) eacts / dissolves in water (1)		[2]
	 6	eaction occurs (1) mmonia is alkaline / neutralisation / hydrogen chloride (1) mmon <u>ium</u> chlor <u>ide</u> formed (1) note: correct equation scores (3)		[3]
	t	ed litmus (1) urns blue (1) Ilow: pH / Universal Indicator (1) urns blue / purple (1)		[2]
2	name blead	ution of chlorine in water ad indicator (1) hes / turns white (1) ot allow: halide test		[2]
	name resul or	ric acid ed indicator (1) (1) arium nitrate (1)		
	white or carbo fizzes	precipitate (1) nate (1)		[2]
	hexe brom	ne ine (water) (1) ourises (1)		[2]
		carbon dioxide (1) / cloudy (1)		[2]

Mark Scheme

Syllabus

Paper

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Page 3		3	Mark Scheme		Paper
			Cambridge IGCSE – October / November 2014	0620	62
3	(a)		ntula (1) not allow: spoon		[1]
	(b)	(i)	sulfuric (1)		[1]
		(ii)	reacts quickly at room temperature (1) allow: heat not needed / reacts anyway		[1]
	(c)	(i)	sulfuric acid / the acid (1)		[1]
		(ii)	solution will be acidic / not neutral / impure salt (1)		[1]
	(d)	(i)	crystals appear / description of using glass rod (1) not: precipitate / evaporate to dryness		[1]
		(ii)	lose water / dehydrate (1) allow: reference to anhydrous ignore: break down of crystals / powder forms		[1]
4	(a)	ten all 6 c 5 c 4 o	ple of results reperature boxes completed correctly (3) 7 correct (3) rect (2) rect (1) refewer correct (0)		[3]
	(b)	all 6 c 5 c 4 o	points correctly plotted (3) 7 correct (3) orrect (2) orrect (1) r fewer correct (0) o intersecting straight line graphs drawn with a ruler (1)		[4]
	(c)	(i)	value from graph, 50 (°C) (1) ± 1 shown clearly (1)		[2]
		(ii)	value from graph, 34 ± 1 (1) unit cm ³ (1) shown clearly (1) note: if tie-line not to peak of graph, max 1, for unit.		[3]
	(d)		dium hydroxide (1) s volume of acid used was greater (1)		[2]

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	(e)	exothermic (1)		[1]
	(f)	room / initial temperature / 26 °C (1) ignore: 20 °C reaction finished owtte (1)		[2]
	(g)	repeat (1) compare results (1) allow: take mean / average (1) ignore: references to insulation		[2]
5	test	s on solution A		
	(a)	yellow / brown / orange (1) allow: combination of above colours do not allow: red, but allow: red-brown		[1]
	(b)	(orange / red) <u>brown</u> (1) allow: rusty precipitate (1)		[2]
	(c)	(orange / red) brown precipitate (1)		[1]
	(d)	white precipitate (1)		[1]
	(i)	aluminium (1) sulfate (1) list principle applies here		[2]
6	(a)	filter solution (1) wash with water (1) dry (1) do not allow: evaporate to dryness		[3]
	(b)	known volume of oven cleaner (1) add named acid (1) with named apparatus (1) indicator (1) observe colour change (1) note volume added (1) repeat with other sample (1) valid comparison (1)		max [6]

Mark Scheme

Syllabus

Paper

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